

VI. *On a General Law of Density in Saturated Vapours.* By J. J. WATERSTON, Esq.
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THE relation between the pressure and temperature of vapours in contact with their generating liquids has been expressed by a variety of empirical formulæ, which, although convenient for practical purposes, do not claim to represent any general law. Some years ago, while examining a mathematical theory of gases, I endeavoured to find out from the French Academy's experiments, if the density of steam in contact with water followed any distinct law with reference to the temperature measured from the zero of gaseous tension. [By RUDBERG's experiments, confirmed by MAGNUS and REGNAULT, this zero is -461° in FAHR. scale, or $-273^{\circ}89$ in the Centigrade scale. *Temperatures reckoned from this zero I shall call G temperatures to save circumlocution.*] If t represents the G temperature, Δ the density of a gas or a vapour, and p its elastic force, the equation

$$t\Delta=p \dots\dots\dots (1.)$$

represents the well-known laws of MARRIOTTE and of DALTON and GAY-LUSSAC. The function that expresses a general relation between p and t in vapours must include a more simple function, expressing a general relation between Δ and t . The proper course, therefore, seemed to be to tabulate the quotients $\frac{p}{t}$ from the experiments of the Academy and to project them into a curve. Now, for reasons connected with the *vis viva* theory of gases, which represents the G temperature as a square quantity, I projected these densities as ordinates to *the square root of the G temperatures as abscissæ*, and the curve traced out was of the parabolic kind, but of high power. To reduce this, because density is a cubic quantity, I tabulated their cube roots and set them off as ordinates to the same abscissæ. The result was gratifying, for the familiar conic parabola made its appearance. To ascertain whether this curve was exactly the conic parabola, I tabulated the square root of these ordinates, corresponding with the sixth root of the densities, and laid them off as new ordinates to the same abscissæ. The result is shown in the accompanying Chart, Plate VII., under the title *French Academy's Steam*. The observations are denoted by dots thus •, and it will be remarked that they range with great precision in a straight line, any slight divergence being sometimes to the right and sometimes to the left; precisely as might be expected from small errors of observation. Other series of experiments on steam were projected in a similar manner; and it was found that although no two exactly agreed with each other, yet that each set ranged in a straight line nearly.

The vapours of ether, alcohol, and sulphuret of carbon were tried in the same way, and found to conform to the same law. I have since added to the Chart M. AVOGADRO's observations on the vapour of mercury, which will be found remarkably in accordance; also Dr. FARADAY's experiments on liquefied gases, given in the Philosophical Transactions for 1845. Of these, olefiant gas (No. 1, p. 160) is remarkably in accordance; also the nitrous oxide (No. 2, p. 168), ammonia, cyanogen, sulphurous acid, and carbonic acid at the upper part of its range. Muriatic acid, sulphuretted and arseniuretted hydrogen do not show the same regularity.

The coordinates of the points being the square root of the G temperature and the sixth root of the densities, the equation to the straight line that passes through the points expresses the sixth root of the density in terms of the square root of the G temperature.

Thus let t_1 = the G temperature corresponding to Δ_1 density of vapour: set off $AC = \sqrt{t_1}$ and $CD = \sqrt[6]{\Delta_1}$. A second observation treated in the same way gives AE, EF , which determines the position of the line of vapour BF . Thus let $AC = g$, and $\cot FBE = h = \frac{DL}{FL} = \left\{ \frac{\sqrt{t_2} - \sqrt{t_1}}{\sqrt[6]{\Delta_2} - \sqrt[6]{\Delta_1}} \right\}$ and

$$g = AB = \sqrt{t_1} - h\sqrt[6]{\Delta_1}.$$

The constants g and h are thus determined from two observations, and the equation for the density Δ at any other G temperature is

$$\left\{ \frac{\sqrt{t} - g}{h} \right\}^6 = \Delta, \quad \dots \dots \dots (2.)$$

and for the pressure,

$$\left\{ \frac{\sqrt{t} - g}{h} \right\}^6 t = p. \quad \dots \dots \dots (3.)$$

The following are the equations for the series of observations given on the Chart; the G temperatures are in degrees of FAHR. scale, and the values given to h are made to give the pressure in inches of mercury.

$$\text{Mercury (AVOGADRO)} \quad \dots \quad p = t \left\{ \frac{\sqrt{t} - 22.606}{20.00} \right\}^6.$$

Oil of turpentine (URE).

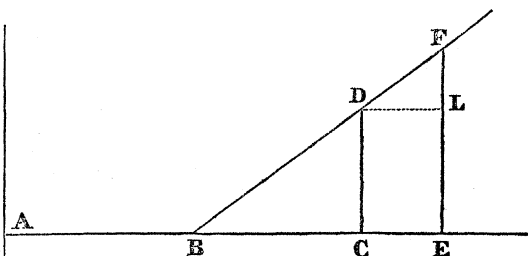
Salt water saturated (WATT).

$$\text{Water (French Academy and SOUTHERN)} \quad \dots \quad p = t \left\{ \frac{\sqrt{t} - 19.492}{10.830} \right\}^6.$$

$$\text{Alcohol, sp. gr. .813 (URE)} \quad \dots \quad p = t \left\{ \frac{\sqrt{t} - 19.287}{9.800} \right\}^6.$$

$$\text{Sulphuret of carbon (MARX)} \quad \dots \quad p = t \left\{ \frac{\sqrt{t} - 16.254}{12.76} \right\}^6.$$

$$\text{Sulphuric ether (DALTON)} \quad \dots \quad p = t \left\{ \frac{\sqrt{t} - 16.860}{10.990} \right\}^6.$$



Sulphurous acid (FARADAY)	$p=t\left\{\frac{\sqrt{t}-14\cdot667}{11\cdot194}\right\}^6.$
Cyanogen (FARADAY)	$p=t\left\{\frac{\sqrt{t}-13\cdot846}{11\cdot542}\right\}^6.$
Ammonia (FARADAY)	$p=t\left\{\frac{\sqrt{t}-13\cdot317}{11\cdot050}\right\}^6.$
Arseniuretted hydrogen (FARADAY)	$p=t\left\{\frac{\sqrt{t}-12\cdot929}{10\cdot264}\right\}^6 (?)$
Sulphuretted hydrogen (FARADAY)	$p=t\left\{\frac{\sqrt{t}-12\cdot957}{9\cdot878}\right\}^6 (?)$
Muriatic acid (FARADAY)	$p=t\left\{\frac{\sqrt{t}-12\cdot060}{9\cdot413}\right\}^6 (?)$
Carbonic acid (FARADAY)	$p=t\left\{\frac{\sqrt{t}-11\cdot997}{8\cdot857}\right\}^6.$
Nitrous oxide (FARADAY), No. 2, p. 168 . .	$p=t\left\{\frac{\sqrt{t}-8\cdot936}{11\cdot604}\right\}^6.$
Olefiant gas (FARADAY), No. 1, p. 160 . .	$p=t\left\{\frac{\sqrt{t}-10\cdot352}{10\cdot152}\right\}^6.$

While projecting different sets of observations of the same vapour, the attention is forcibly drawn to the tendency which the line shows to alter its inclination on the most trifling change of circumstances. This inclination depends on the value of the denominator h . It will be remarked that this constant is nearly the same for the vapours of ether and of water, also for a considerable number of the liquified gases. In fact the difference is not greater than might be caused by a slight want of purity. Even the specific nature of the vessel in which the observations are made seems materially to affect the results. Dr. FARADAY employed tubes of from one-fifth to one-seventh inch bore. I believe it will be found that the density is considerably less in these tubes than in larger vessels; I have had proof of this in the case of ether heated in such tubes to 280° . The capillary attraction of the sides even at high pressures seems to affect the whole surface of the liquid, and to check the formation of vapour. We must not forget also, that Dr. FARADAY's temperatures are confessedly not so low as a correct thermometer would have indicated. The liquid he employed to measure the temperature was alcohol, and the degrees were graduated with the same capacity as between 32° and 212° . Now from the experiments that have been made on the expansion of alcohol by M. GAY-LUSSAC, and more recently by M. MUNCKE, it appears that the mean contraction between 32° and 212° is 6·2 thousandths of the volume at 173° , its boiling-point, and that this amount diminishes very sensibly with the temperature; and if it continued to do so at the same rate from 32° to -100° , we should have Dr. FARADAY's -100° equal to less than -200° by the air-thermometer. It is clear that until the value of these degrees are known with reference to the air-thermometer, the importance of Dr. FARADAY's observations cannot be fully estimated.

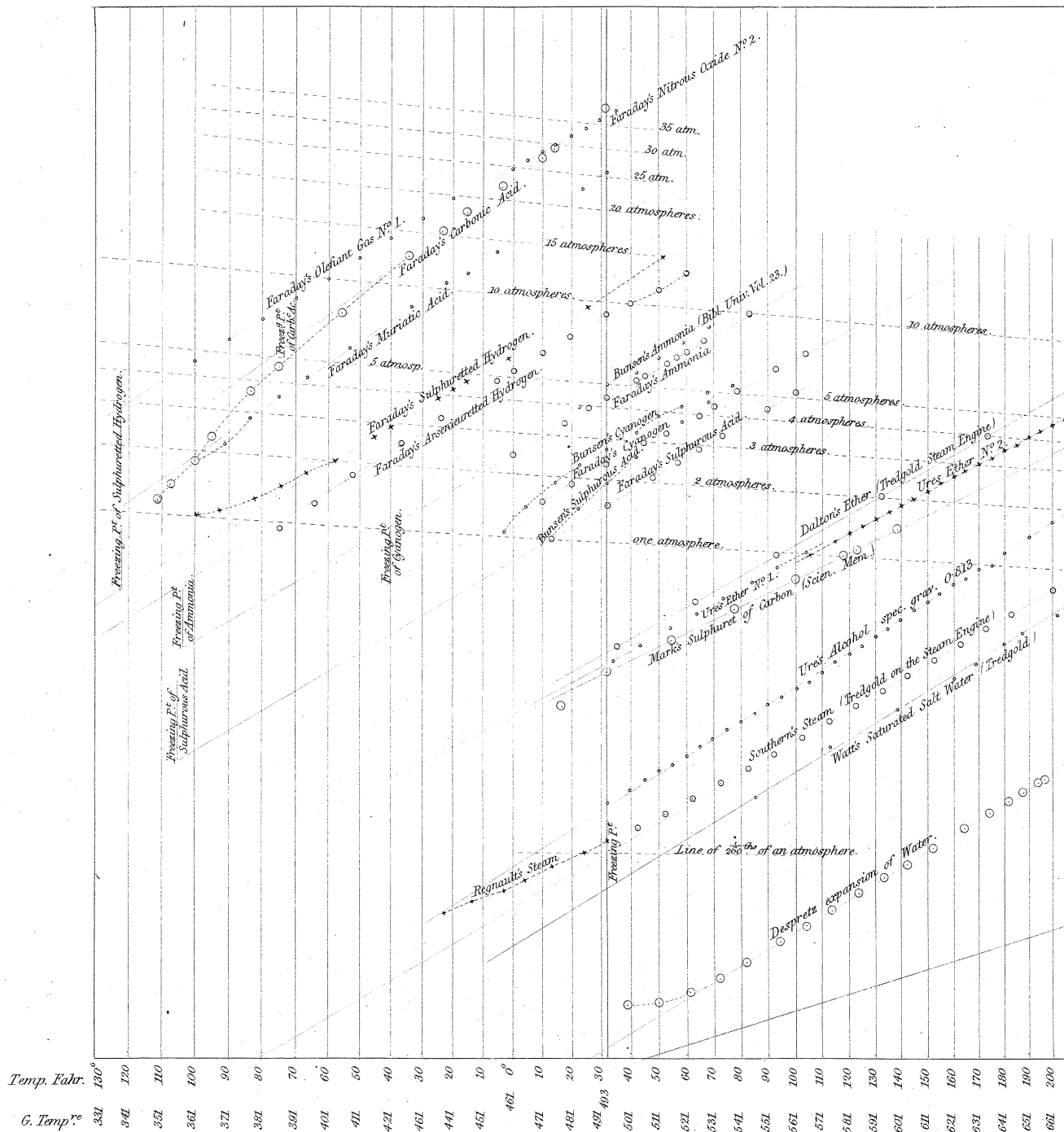


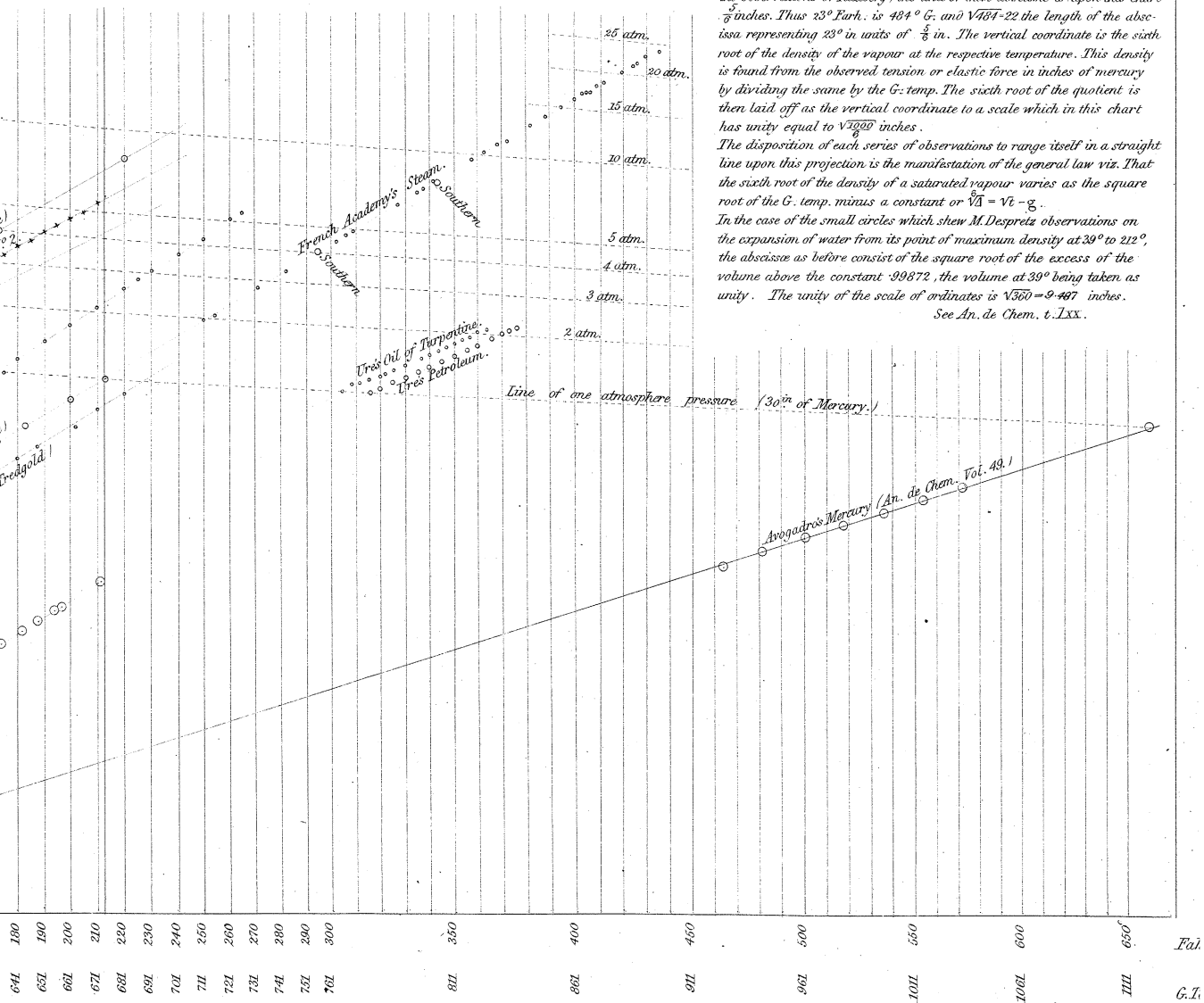
Chart of Saturated Vapours illustrating the general law of their density.

Explanation. The observations are severally marked by small circles, (o) dots, (•) or crosses, (+) and the position of each on the Chart is fixed by coordinates as follows. The horizontal coordinate or abscissa is the square root of the G. temperature whose zero is situated at 461° Fahr. (being the zero of gaseous tension according to the observations of Rudberg) the unit of these abscissae is upon this Chart $\frac{5}{8}$ inches. Thus 23° Fahr. is 484° G. and $\sqrt{484}=22$ the length of the abscissa representing 23° in units of $\frac{5}{8}$ in. The vertical coordinate is the sixth root of the density of the vapour at the respective temperature. This density is found from the observed tension or elastic force in inches of mercury by dividing the same by the G. temp. The sixth root of the quotient is then laid off as the vertical coordinate to a scale which in this chart has unity equal to $\sqrt[6]{\frac{1000}{9}}$ inches.

The disposition of each series of observations to range itself in a straight line upon this projection is the manifestation of the general law viz. That the sixth root of the density of a saturated vapour varies as the square root of the G. temp. minus a constant or $\sqrt{A} = \sqrt{t - c}$.

In the case of the small circles which shew M. Despretz observations on the expansion of water from its point of maximum density at 39° to 212°, the abscissae as before consist of the square root of the excess of the volume above the constant .99872, the volume at 39° being taken as unity. The unity of the scale of ordinates is $\sqrt[6]{\frac{1000}{9}} = 9.487$ inches.

See An. de Chem. t. LXX.



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Chart of Saturated Vapours

illustrating the general law of their density.

Explanation. The observations are generally marked by small circles, (o) dots, (+) or crosses, (x) and the position of each on the chart is fixed by coordinates as follows. The horizontal coordinate or abscissa is the square root of the G , temperature whose zero is at -461° Fahr. (being the zero of gaseous tension according to the observations of Rudberg) the unit of these abscissas is upon this Chart $\frac{1}{2}$ inch. Thus 32° Fahr. is 461° G., and $\sqrt{461-22}$ the length of the abscissa representing 32° in units of $\frac{1}{2}$ in. The vertical coordinate is the sixth root of the density of the vapour at the respective temperature. This density is found from the observed tension or elastic force in inches of mercury by dividing the same by the G . temp. The sixth root of the quotient is then laid off as the vertical coordinate to a scale which in this chart has unity equal to $\sqrt[6]{1000}$ inches.

The disposition of each series of observations to range itself in a straight line upon this projection is the manifestation of the general law viz. That the sixth root of the density of a saturated vapour varies as the square root of the G . temp. minus a constant or $\sqrt{A - V} - g$. In the case of the small circles which show M. Despretz's observations on the expansion of water from its point of maximum density at 39° to 32° , the abscissa as before consist of the square root of the excess of the volume above the constant 99872, the volume at 39° being taken as unity. The unity of the scale of ordinates is $\sqrt[6]{1000} = 9.487$ inches.

See An. de Chem. t. LXX.

